

## Influence of Quaternization on UCST Properties of Hydroxyl-Derivatized Polymers

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A series of hydroxyl-derivatized quaternized polymers were successfully synthesized by atom transfer radical polymerization (ATRP) and Cu(I)-catalyzed 1,3-dipolar cycloaddition of azide and alkynes (click chemistry), followed by quaternization reactions. ATRP was employed to synthesize poly(2-hydroxyethyl methacrylate) (PHEMA), followed by introduction of alkyne groups using pentynoic acid, leading to HEMA-Alkyne. 2-Azido-1-ethanol and 3-azido-1-propanol were combined with the HEMA-Alkyne backbone *via* click reaction, resulting in triazole-ring containing hydroxyl-derivatized polymers. Quaternization reactions with methyl iodide were conducted on the triazole ring of each polymer. Molecular weight, molecular weight distribution, and the degree of quaternization (DQ) were determined by gel permeation chromatography (GPC) and  $^1\text{H}$  NMR spectroscopy. The average molecular weight ( $M_n$ ) of the resulting polymers ranged from  $5.9 \times 10^4$  to  $1.05 \times 10^5$  g/mol depending on the molecular architecture. The molecular weight distribution was low ( $M_w/M_n = 1.26$ - $1.38$ ). The transmission spectra of the 0.1 wt % aqueous solutions of the resulting quaternized polymers at 650 nm were measured as a function of temperature. Results showed that the upper critical solution temperature (UCST) could be finely controlled by the level of DQ.

**Key Words :** ATRP, Quaternized, Click reactions, UCST, Thermoresponsive

### Introduction

Stimuli-responsive polymers undergo dramatic changes of their physical and chemical properties in response to small external changes in the environmental condition, including pH, temperature, and light, *etc.*<sup>1-6</sup> Especially, thermoresponsive polymers exhibit thermally induced, reversible phase transitions that are caused by a sudden change in the solvation state.<sup>7</sup> Depending on whether reversible phase transition is found at low or high temperatures, an upper or lower critical solution temperature (UCST<sup>8-10</sup> or LCST) exists, respectively. LCST systems are usually restricted to an aqueous solvent environment, which are primarily of interest for biomedical applications. UCST systems, however, have been also studied in organic solvents.<sup>11-13</sup>

UCST behavior can happen when interpolymer interactions are gradually weakened upon heating. For example, zwitterionic polybetaines have strong electrostatic (ionic) interactions between polymer chains. These interaction can be readily disrupted upon heating, which renders polymer to be soluble in water. The UCST of polymers is also driven by the presence of ionic impurities or hydrogen bonding<sup>14</sup> between hydrogen bond donors and acceptors such as poly(allylurea),<sup>15</sup> poly(6-(acryloyloxymethyl)uracil),<sup>16</sup> and poly(*N*-acryloylasparaginamide (NAAAM)).<sup>17</sup>

Previously, we reported the synthesis and UCST behavior of a series of ureido-derivatized polymers with different degrees of quaternization (DQ) of their triazole rings.<sup>18</sup> As a continuous effort to develop polymeric UCST systems in water, we synthesized thermoresponsive hydroxyl-derivatized quaternized polymers. This work represents the first example of UCST tuning of hydroxyl-derivatized polymers

simply by DQ.

### Experimental

**Materials.** 2-Hydroxyethyl methacrylate (HEMA, 95%, Tokyo Chemical Industry; TCI) passed through basic alumina column to remove the monomethyl ether hydroquinone inhibitor. 2,2'-Bipyridyl (bpy), ethyl  $\alpha$ -bromoisobutyrate (EBiB), CuCl (99%), CuCl<sub>2</sub> (99%), CuBr (98%), 4-dimethylamino-pyridine (DMAP), *N,N'*-dicyclohexylcarbodiimide (DCC), 4-pentynoic acid, diethyl ether, *N,N,N',N'',N'''*-pentamethyldiethylenetriamine (PMDETA), 2-bromoethanol, 3-chloro-1-propanol, and methyl iodide were purchased from Aldrich with the highest purity and were used as received without further purification. Sodium azide (NaN<sub>3</sub>), anisole, methanol, and methylene chloride (MC) were purchased from TCI and used as received.

**Instrumentation.**  $^1\text{H}$  NMR spectra were recorded in DMSO-*d*<sub>6</sub> and CDCl<sub>3</sub> on a Bruker avance 300 MHz NMR spectrometer. The apparent molecular weights and molecular weight distributions were measured by gel permeation chromatography (GPC) (Agilent Technologies 1200 series) using a polystyrene (PS) standard, with DMF as the eluent at 30 °C and a flow rate of 1.00 mL/min. The cloud points of the polymer solutions in water were measured on a Thermo Scientific Evolution 600 UV-Vis spectrophotometer equipped with a digital temperature controller. A 650 nm wavelength was used to determine UCST. Sample cells (1 cm) were thermostated with a water jacket ranged from 15 to 95 °C with a cooling ingrate of 1 °C/min. The cloud point was defined as the middle point of the transmittance change.

**2-Azido-1-ethanol (1):** NaN<sub>3</sub> (16.8 g, 0.258 mol) was

added to a solution of 2-bromoethanol (10 g, 0.086 mol) in water (100 mL). The reaction mixture was heated to 60 °C for 15 h, extracted with diethyl ether, dried over MgSO<sub>4</sub>, and concentrated to give volatile colorless oil. <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>) δ 3.72 (2H, t, CH<sub>2</sub>CH<sub>2</sub>OH); 3.39 (2H, t, N<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>); 2.80 (1H, s, CH<sub>2</sub>OH).

**3-Azido-1-propanol (2):** <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>) δ 3.68 (2H, t, CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>OH); 3.39 (2H, t, N<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>); 2.730 (1H, s, CH<sub>2</sub>OH); 1.78 (2H, m, CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>).

**PHEMA.** PHEMA was synthesized as described previously.<sup>19</sup> M<sub>n</sub> = 59,300 g/mol, M<sub>w</sub>/M<sub>n</sub> = 1.34. <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>) δ 5.0 (1H, s, CH<sub>2</sub>OH); 4.1 (2H, s, O-CH<sub>2</sub>-CH<sub>2</sub>); 3.7 (2H, s, O-CH<sub>2</sub>-CH<sub>2</sub>); 2.0-1.9 (2H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)); 1.1-0.94 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)).

**PHEMA-Alkyne.** PHEMA-Alkyne was synthesized as described previously.<sup>19</sup> M<sub>n</sub> = 69,500 g/mol, M<sub>w</sub>/M<sub>n</sub> = 1.26. <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>) δ 4.4-4.27 (4H, d, -O-CH<sub>2</sub>-CH<sub>2</sub>-O-); 2.69-2.68 (2H, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-); 2.58-2.56 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.18-1.74 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>), CH<sub>2</sub>-C-CH); 1.2-0.8 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)).

**PEA (PHEMA-Alkyne-click-1).** The ratio of reagent [PHEMA-Alkyne]<sub>0</sub>/[2-Azido-1-ethanol]<sub>0</sub>/[CuBr]<sub>0</sub>/[PMDETA]<sub>0</sub> was 1/2/0.1/0.1. The click reaction between PHEMA-Alkyne (0.5 g, 2.38 mmol) and 2-azido-1-ethanol (0.410 g, 4.76 mmol) was conducted in 5 mL of DMF using CuBr/PMDETA as a catalyst. After 10 h, the polymer solution was exposed to air, diluted with DMF, and passed through neutral alumina to remove the copper catalyst. The resulting polymer was precipitated in diethyl ether and dried in a vacuum oven for 24 h. M<sub>n</sub> = 97,700 g/mol, M<sub>w</sub>/M<sub>n</sub> = 1.31. <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>) δ 7.77 (1H, s, triazole); 4.99 (1H, t, CH<sub>2</sub>CH<sub>2</sub>OH); 4.31 (2H, t, triazol-CH<sub>2</sub>-CH<sub>2</sub>-OH); 4.16-4.07 (4H, d, -O-CH<sub>2</sub>-CH<sub>2</sub>-O-); 3.74 (2H, t, triazol-CH<sub>2</sub>-CH<sub>2</sub>-OH); 2.88 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.65 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.05-1.3 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>), CH<sub>2</sub>-C-CH); 0.92-0.77 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)).

**PPA (PHEMA-Alkyne-click-2).** M<sub>n</sub> = 105,000 g/mol, M<sub>w</sub>/M<sub>n</sub> = 1.38. <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>) δ 7.77 (1H, s, triazole); 4.60 (1H, t, triazol-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>2</sub>-OH); 4.32 (2H, t, triazol-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>2</sub>-OH); 4.07-4.02 (4H, d, -O-CH<sub>2</sub>-CH<sub>2</sub>-O-); 3.43 (2H, t, triazol-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>2</sub>-OH); 2.85 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.64 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.05-1.3 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)CH<sub>2</sub>-C-CH); 3.43 (2H, m, triazol-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>2</sub>-OH); 0.92-0.77 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)).

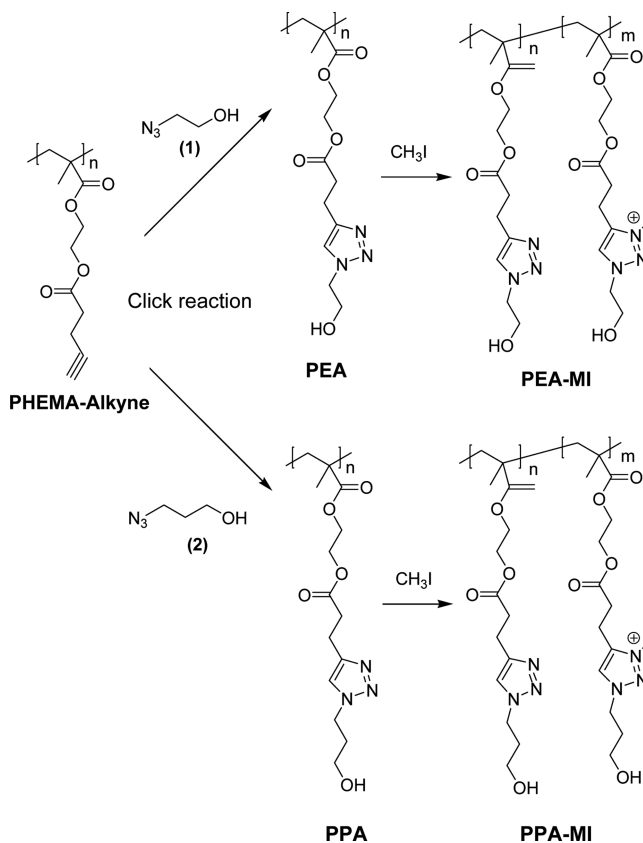
**Quaternization of PEA (PEA-MI).** A typical procedure for quaternization of PEA with methyl iodide (1:1.5 molar ratio) was as follows. A solution of PEA (0.0377 g, 0.2 mmol of triazole groups) and methyl iodide (0.043 g; 0.3 mmol) in DMF (0.5 mL) was stirred for 16 h at 45 °C. Quaternized PEA was precipitated in DMF, filtered, and dried under vacuum. Reactions with three different molar ratios (1:1.5, 1:2.0, 1:2.5) of PEA to methyl iodide were carried out, and the degree of quaternization (DQ) was 4%, 7%, and 10%, respectively. <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>) δ 8.75 (1H, s, quaternized triazole); 7.77 (1H, s, triazole); 4.99 (1H, t, CH<sub>2</sub>CH<sub>2</sub>OH); 4.61 (3H, s, triazole-CH<sub>3</sub>); 4.31

(2H, t, triazole-CH<sub>2</sub>-CH<sub>2</sub>-OH); 4.17-4.07 (4H, d, -O-CH<sub>2</sub>-CH<sub>2</sub>-O-); 3.74 (2H, t, triazol-CH<sub>2</sub>-CH<sub>2</sub>-OH); 2.88 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.65 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.05-1.3 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>), CH<sub>2</sub>-C-CH); 0.92-0.77 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)).

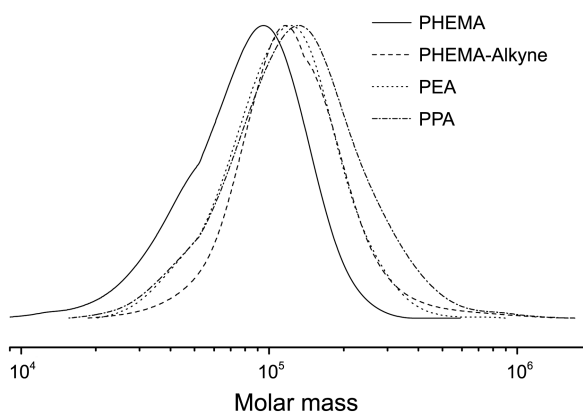
**Quaternization of PPA (PPA-MI).** Reactions with three different molar ratios (1:1.5, 1:2.0, 1:2.5) of PPA to methyl iodide were carried out, and the degree of quaternization (DQ) was 3%, 8%, and 11%, respectively. <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>) δ 8.75(1H, s, triazole); 7.77 (1H, s, triazole); 4.60 (1H, t, triazol-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>2</sub>-OH); 4.32 (2H, t, triazole-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>2</sub>-OH); 4.07-4.02 (4H, d, -O-CH<sub>2</sub>-CH<sub>2</sub>-O-); 3.43 (2H, t, triazol-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>2</sub>-OH); 2.85 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.64 (2H, s, d, O-(C=O)-CH<sub>2</sub>-CH<sub>2</sub>-C-); 2.05-1.3 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)CH<sub>2</sub>-C-CH); 3.43 (2H, m, triazol-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>2</sub>-OH); 0.92-0.77 (3H, d, CH<sub>2</sub>-C(CH<sub>3</sub>)).

## Results and Discussion

The strategy employed in this study is schematically illustrated in Scheme 1. PHEMA with a controlled molecular weight and low polydispersity was synthesized *via* atom transfer radical polymerization (ATRP)<sup>20,21</sup> from EBiB initiator using CuCl/bpy as the catalyst. The molecular weight and molecular weight distribution of the resulting PHEMA were obtained using a GPC DMF line with PS standards (M<sub>n</sub> = 59 300 g/mol, M<sub>w</sub>/M<sub>n</sub> = 1.34) (Figure 1 and Table 1). The



**Scheme 1.** Synthesis of hydroxyl-derivatized quaternized polymers by combination of ATRP and click chemistry, followed by quaternization with methyl iodide.



**Figure 1.** DMF GPC traces of PHEMA, PHEMA-Alkyne, PEA, and PPA.

**Table 1.** Summary of DMF GPC results

	$M_n, \text{theory}^a$ (g/mol)	$M_n, \text{app}^b$ (g/mol)	PDI <sup>b</sup>
PHEMA	26 000 <sup>a</sup>	59 300	1.34
PHEMA-alkyne	31 500	69 500	1.26
PEA	39 900	97 700	1.31
PPA	38 600	105 000	1.38

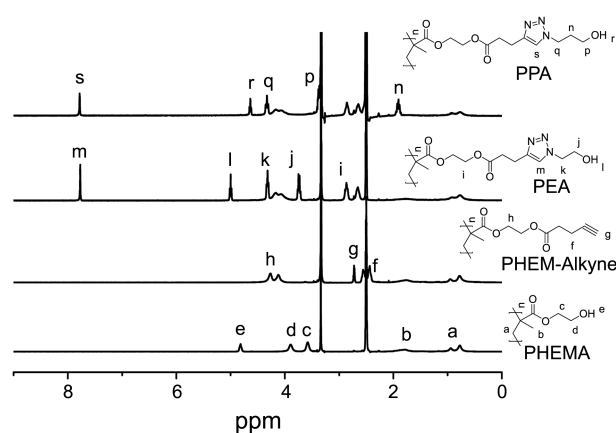
<sup>a</sup>Theoretical molecular weight determined from monomer conversions.

<sup>b</sup>Apparent number-average molecular weight and polydispersity index (PDI) determined by DMF GPC with PS calibration.

apparent molecular weight obtained by GPC was higher than the theoretical molecular weight ( $M_n, \text{theory} = 26\,000\text{ g/mol}$ ) calculated from the HEMA conversion ( $M_n, \text{theory} = \text{conversion} \times \text{MW}_{\text{HEMA}} \times [\text{HEMA}]_0/[\text{EBiB}]_0$ ) due to the differences in hydrodynamic volumes of PHEMA and PS standards. In the next step, DCC coupling of PHEMA with 4-pentynoic acid yielded PHEMA-Alkyne ( $M_n = 69\,500\text{ g/mol}$ ,  $M_w/M_n = 1.26$ ) as the polymer backbone. The successful synthesis of PHEMA-Alkyne was confirmed by  $^1\text{H}$  NMR spectra. It was monitored that peak (e) of PHEMA at 4.85 ppm disappeared while peak (g) of PHEMA-Alkyne appeared at 2.7 ppm (Figure 2). The GPC traces in Figure 1 show that the apparent molecular weight of PHEMA-Alkyne backbone increased slightly after transformation.

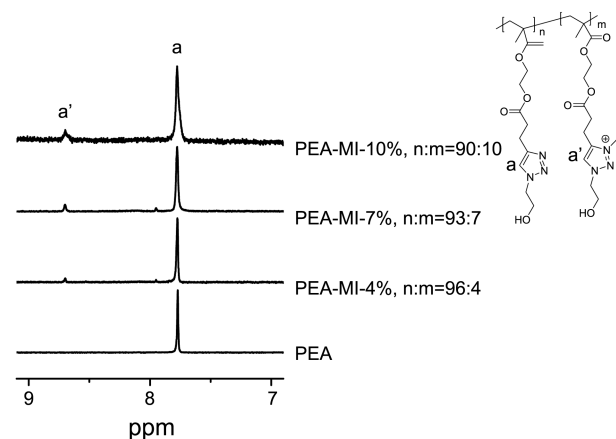
2-Azido-1-ethanol (1) and 3-azido-1-propanol (2) were successfully added to the PHEMA-Alkyne backbone in DMF with a CuBr/PMDETA complex at room temperature by click chemistry,<sup>22,23</sup> resulting in PEA and PPA, respectively. To ensure 100% conversion, the twice molar excess amount of 2-azido-1-ethanol to PHEMA-Alkyne backbone was used. The quantitative conversions were observed by  $^1\text{H}$  NMR spectra, which provided evidence of the successful synthesis of the hydroxyl-functionalized polymers (Figure 2). After the click reaction, new peaks (m and l) representing the proton on the 1,2,3-triazole rings appeared at 7.9 ppm and the proton of terminal  $-\text{OH}$  at 5.0 ppm appeared. The integral ratio between peak (m) and (l) was exactly 1, indicating the quantitative click reaction.

In the final polymer modification step, PEA and PPA were

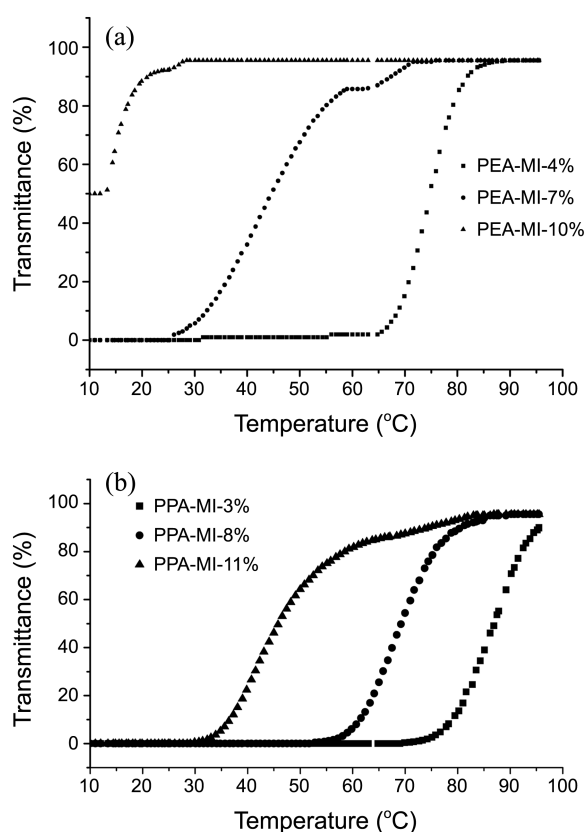


**Figure 2.**  $^1\text{H}$  NMR spectra of PHEMA, PHEMA-Alkyne, PEA, and PPA.

partially quaternized with different amounts of methyl iodide to yield various levels of the quaternized polymers, PEA-MI and PPA-MI. Reactions under three different molar feed ratios (1:1.5, 1:2.0, and 1:2.5) of PEA to methyl iodide were conducted in DMF at 45 °C. The degree of quaternization (DQ) for the PEA series was 4%, 7%, and 10%, respectively, which was estimated by  $^1\text{H}$  NMR spectra (Figure 3). As described in our previous results,<sup>18</sup> the prediction of DQ is difficult due to the quaternization reaction being relatively inefficient. For the quaternization reaction of PEA, a two-fold molar excess of methyl iodide over triazole groups in the repeating units of PEA resulted in only a 7% DQ (PEA-MI-7%). As shown in Figure 3, a new peak (a') representing the proton on the quaternized triazole ring appeared at 8.65 ppm, while the intensity of the proton on the triazole ring without quaternization (a) decreased. As DQ increased, the intensity of peak (a') increased and that of peak (a) decreased. DQ of PEA was calculated from the relative integration areas of peak (a) and (a'). DQ of PPA was 3%, 8%, and 11%, which was calculated similarly as described above. Due to volatile nature of methyl iodide (boiling point = 42–43 °C) and relatively high reaction temperature (45 °C), loss of methyl iodide is inevitable during the reaction, which led to



**Figure 3.**  $^1\text{H}$  NMR spectra of PEA and its quaternized polymers with different DQ (PEA-MI-4%, PEA-MI-7%, and PEA-MI-10%).



**Figure 4.** Effect of quaternization on UCST transitions of 1.0 % aqueous solution: (a) PEA series and (b) PPA series measured by percent transmission at 650 nm.

inefficient quaternization.

The UCST transitions of polymers in dilute aqueous solution were monitored by observing changes in the percent transmission at a fixed wavelength (650 nm) during cooling cycles *via* UV-Vis spectroscopy. The concentration of all aqueous solutions was fixed at 1.0 wt % because it is relatively easy to detect cloud points at higher concentrations. The cloud point was defined as the middle point of the transmittance change. Note that quaternized PEA and PPA were not soluble in water even at elevated temperature, meaning that UCST values of these polymers are higher than 100 °C. If these polymers become more hydrophilic by quaternization, one can expect that the UCST would gradually decrease. Indeed, the hydrophilicity of PEA and PPA increased with increasing DQ, which caused the UCST to decrease. It was found that UCST values for the quaternized series of PEA decreased from 75 to 15 °C as the DQ increased from 4 to 10% (Figure 4(a)). Similar trends were observed for the quaternized series of PEA (Figure 4(b)). UCST values decreased from 88 to 48 °C as the DQ increased from 4 to 10%. Since PPA is more hydrophobic than PEA, UCST values of quaternized PEA are higher than those of quaternized PPA when similar DQ was achieved.

## Conclusions

We demonstrated the synthesis of thermoresponsive hydroxyl-derivatized quaternized polymers with UCST properties *via* a combination of ATRP and Cu(I)-catalyzed 1,3-dipolar cycloaddition of azide and alkynes, followed by quaternization reaction. The successful synthesis of these polymers was verified by <sup>1</sup>H NMR spectroscopy and GPC. It was demonstrated that the polymer UCST could be affected by the existence of ionic groups induced by quaternization of methyl iodide as well as chain-end hydroxyl groups. We also demonstrated that UCST values of hydroxyl-derivatized quaternized polymers were tuned simply by the DQ of the triazole rings.

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