

Synthesis of a Di-*N*-cyanoethylated Tetraaza Macrocycle Containing Eight *C*-Methyl Groups and Its Nickel(II) Complex: Effects of the Methyl Groups on Their Properties

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A new di-*N*-cyanoethylated 14-membered tetraaza macrocycle 1,8-bis(2-cyanoethyl)-3,5,7,7,10,12,14,14-octamethyl-1,4,8,11-tetraazacyclotetradecane (L^2) and its nickel(II) complex $[NiL^2(OAc)]^-$ have been prepared. The square-planar complex $[NiL^2](ClO_4)_2$ can be prepared by addition of $HClO_4$ to a hot aqueous solution of $[NiL^2(OAc)]^-$. The Ni-N (tertiary amino group) bond distances (2.008 Å) of $[NiL^2](ClO_4)_2$ are relatively long, and the complex exhibits a *d-d* transition band at unusually long wavelength (*ca.* 515 nm). The complex $[NiL^2](ClO_4)_2$ rapidly reacts with acetate ion or ethylenediamine (en) to produce $[NiL^2(OAc)]^+$ or $[Ni(en)_3]^{2+}$, respectively, and is readily decomposed in NaOH (0.01 M) solution. The chemical properties of $[NiL^2]^{2+}$ as well as its synthetic procedure are quite different from those for other related 14-membered tetraaza macrocyclic complexes. Effects of the *N*-cyanoethyl and *C*-methyl groups on the properties of L^2

Keywords: Macrocyclic compounds. Nickel(II) complexes. Cyanoethyl pendant arms. Crystal structure.

Introduction

There has been considerable interest in the preparation of various 14-membered tetraaza macrocyclic compounds containing additional *N*- and/or *C*-alkyl groups because of the effects of the alkyl groups on their chemical properties and coordination geometry.¹⁻¹⁶ In most cases, the effects of the *C*-alkyl groups are relatively weak or negligible and have received much less attention than those of the *N*-alkyl groups. However, some recent works indicate that coordination geometry and/or chemical properties of such compounds can be also affected significantly by the *C*-alkyl groups if there are additional *N*-substituents.⁶⁻¹⁰

Some 14-membered tetraaza macrocycles containing *N*-cyanoethyl pendant arms, such as L^7 , L^8 , L^{10} , and L^{11} , have been prepared.¹⁰⁻¹³ It has been observed that chemical reactivity of such compounds are strongly affected by the *C*-substituents. For example, although $[NiL^7]^{2+}$ can be prepared readily by the reaction of the metal ion with the macrocycle,^{11,12} the reaction of L^{10} with Ni^{2+} produces the di-*N*-cyanoethylated complex $[NiL^{10}]^{2+}$,¹⁰ the complex of L^{10} can not be prepared by the reaction.¹⁰ The cyanoethyl pendant arms ($-CH_2CH_2CN$) of $[NiL^7]^{2+}$ are slowly hydrolyzed to the corresponding amide groups ($-CH_2CH_2CONH_2$) in boiling aqueous solutions.^{11,12} On the other hand, $[NiL^{11}]^{2+}$ is completely decyanoethylated to $[NiL^9]^{2+}$ before hydrolysis under similar conditions.¹⁰ Until now, the effects of *C*-substituents on the properties of *N*-cyanoethylated macrocyclic compounds are not thoroughly investigated because examples of such compounds are relatively rare.¹⁰⁻¹² This prompted to us to prepare the di-*N*-cyanoethylated macrocycle L^2 , in which eight methyl groups are attached to the carbon atoms, and its nickel(II) complexes. The nickel(II) complex of L^2 contain-

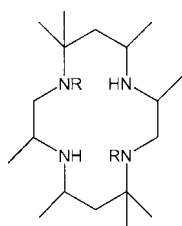
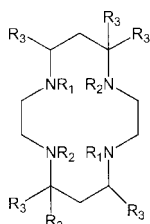
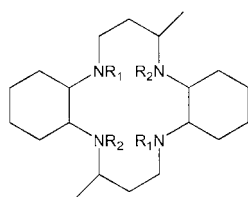
ing six *C*-methyl groups was also prepared for comparison. The major aim of this work was to understand the effects of the *C*-alkyl groups on the reactivity and coordination behavior of the *N*-cyanoethylated macrocyclic compounds. Somewhat surprisingly, it was found that the synthetic procedure and reactivity of the nickel(II) complex $[NiL^2](ClO_4)_2$ as well as the ligand field strength are considerably different from those for the complexes of L^3 , L^4 , L^8 , and L^{11} . Synthesis and properties of L^2 and its nickel(II) complexes are reported, together with the crystal structure of $[NiL^2](ClO_4)_2$.

Experimental Section

Measurements. Infrared spectra were recorded with a Shimadzu IR-440 spectrophotometer, electronic absorption spectra with a Shimadzu UV-160 spectrophotometer, NMR spectra with a Varian Mercury 300 NMR spectrometer, and conductance measurements with a Metrohm Herisau Conductometer E518. Elemental analyses and mass spectral analysis were performed at the Korea Basic Science Institute, Taegu, Korea. Magnetic moments were calculated from magnetic susceptibility data obtained using a Johnson Matthey MK-1 magnetic susceptibility balance.

Synthesis. *Safety Note:* Perchlorate salts of transition metal complexes with organic ligands are often explosive and should be handled with caution.

L^2 . The macrocycle L^1 was prepared according to the reported method.^{15,16} To a methanol solution (20 mL) of L^1 (3.0 g, 9.7×10^{-3} mol) was added an excess of acrylonitrile (2.4 mL, 7.5×10^{-2} mol). The mixture was stirred for *ca.* 30 h at room temperature. The resulting solution was evaporated at room temperature to produce the white solid. The

L¹: R = HL²: R = CH₂CH₂CNL³: R = CH₂CH=CH₂L⁴: R = CH₂CH₂CH₃L⁵: R₁ = R₂ = R₃ = HL⁶: R₁ = R₂ = H; R₃ = MeL⁷: R₁ = R₂ = CH₂CH₂CN; R₃ = HL⁸: R₁ = CH₂CH₂CN; R₂ = H; R₃ = MeL⁹: R₁ = R₂ = HL¹⁰: R₁ = R₂ = CH₂CH₂CNL¹¹: R₁ = CH₂CH₂CN; R₂ = H

solid was filtered, washed with diethyl ether, and dried in air. Yield: ~80%. Anal. Calcd. for C₂₄H₄₆N₆: C, 68.85; H, 11.07, N, 20.07%. Found: C, 68.63; H, 11.17; N, 20.17%. Mass (*m/z*): 419 (M⁺). IR (cm⁻¹): 3235 (νN-H) and 2250 (νCN). ¹³C-NMR (CDCl₃): δ 18.8 (Me), 20.8 (Me), 21.0 (Me), 22.8 (Me), 28.0, 45.5, 46.5, 47.3, 47.7, 56.9, 58.4, and 120.7 (CN) ppm.

[NiL²(OAc)]ClO₄. A methanol solution (20 mL) of Ni(OAc)₂·4H₂O (2.0 g, 8.0 mmol) and L² (2.0 g, 4.8 mmol) was refluxed for *ca.* 1 hr. An excess of NaClO₄ was added to the solution at room temperature, then a green solid was precipitated. The product was filtered, washed with methanol, and recrystallized from a 1.0 × 10⁻³ M NaOAc/MeCN-water (2 : 1) mixture. Yield: ~80%. Anal. Calcd for C₂₆H₄₉N₆NiCl₂O₈: C, 49.11; H, 7.77; N, 13.22%. Found: C, 50.05; H, 8.23; N, 13.54%. IR (cm⁻¹): 3225 (νN-H), 3165 (νN-H), 2250 (νCN), and 1600 (νCO). FAB mass (*m/z*): 535 ([M-ClO₄]⁺) and 476 ([M-ClO₄-CH₃COOH]⁺). Magnetic moment (*μ*_{eff}): 2.65 μ_B at 20 °C.

[NiL²](ClO₄)₂. To a warm acetonitrile solution (10 mL) of [NiL²(OAc)]ClO₄ (0.5 g) was added water (20 mL). The volume of the solution was reduced to *ca.* 20 mL on a hot water bath (70-80 °C), and pH of the resulting solution was adjusted to *ca.* 3.0 by dropping 1.0 M HClO₄. The solution, which was initially green, turned intense orange-red. After the resulting solution had been heated on the water bath for *ca.* 5 min, concentrated HClO₄ (2-3 drops) and NaClO₄ (excess) were added at room temperature to produce the red crystal. The product was filtered, washed with methanol, and

recrystallized from DMSO-water (1 : 1) mixture. Yield: ~60%. Anal. Calcd for C₂₄H₄₆N₆NiCl₂O₈: C, 42.63; H, 6.86; N, 12.43%. Found: C, 42.29; H, 6.93; N, 12.28%. IR (cm⁻¹): 3140 (νN-H) and 2250 (νCN).

[H₂L²](ClO₄)₂. The addition of an excess HClO₄ to a acetonitrile-water solution of [NiL²(OAc)]ClO₄ at room temperature produced a white solid. The product was filtered, washed with methanol, and dried in air. Anal. Calcd for C₂₄H₄₈N₆Cl₂O₈: C, 46.68; H, 7.51; N, 13.61%. Found: C, 45.94; H, 7.56; N, 13.55%. The compound was also obtained by the addition of HClO₄ to the reaction mixture of L² and NiCl₂·6H₂O in methanol.

[NiL⁸](ClO₄)₂. The macrocycle L⁸ was prepared according to the reported method.¹³ A methanol solution (20 mL) of Ni(OAc)₂·4H₂O (1.3 g, 5.1 mmol) and L⁸ (1.5 g, 3.8 mmol) was refluxed for *ca.* 30 min. An excess of HClO₄ was added to the solution, and then the orange-red solid was precipitated. The product was filtered, washed with methanol, and recrystallized from acetonitrile-water (2 : 1) mixture. Yield: ~80%. Anal. Calcd for C₂₂H₄₃N₆NiCl₂O₈: C, 40.76; H, 6.53; N, 12.96%. Found: C, 40.89; H, 6.78; N, 13.07%. IR (cm⁻¹): 3165 (νN-H) and 2250 (νCN).

Crystal Structure Determination. A red crystal of [NiL²](ClO₄)₂, obtained from DMSO/H₂O solution, was mounted on a thin glass fiber. Intensity data were collected on an Enraf-Nonius CAD4 Diffractometer. Unit cell parameters were determined from a least-squares fit of 25 accurately centered reflections (20 < 2θ < 32). Data were collected at 293 K and were corrected for Lorentz and polarization effects. Empirical ψ absorption correction was applied. The SHELXS-86 program was utilized for the heavy atom method.¹⁷ The structure refinements were performed with the SHELXL-93 program on F² data.¹⁸ Heavy atoms (Ni and Cl) were located by use of heavy atom method.

Table 1. Crystal and Refinement Data for [NiL²](ClO₄)₂

Formula	C ₂₄ H ₄₆ Cl ₂ N ₆ NiO ₈
M	676.28
Cryst size (mm)	0.4 × 0.4 × 0.3
Space group	P2 ₁ /n
<i>a</i> (Å)	9.571 (2)
<i>b</i> (Å)	11.079 (2)
<i>c</i> (Å)	14.232 (3)
β (°)	92.14 (2)
<i>V</i> (Å ³)	1508.1 (5)
<i>Z</i>	2
<i>D_c</i> (g cm ⁻³)	1.489
<i>μ</i> / cm ⁻¹	8.76
Range of transmission factor	84.51-99.98%
Scan mode	ω-2θ
No. of unique reflections	2651
No. of observed reflections	2215 (<i>I</i> _o > 2σ(<i>I</i> _o))
2θ max / °	50
No. of parameters refined	187
R1, wR2 ^a , GOF	0.0370, 0.1005, 1.05

^a ω = 1/[σ²(*F_o*²) - (0.0565*P*)² + 1.8617*P*] where *P* = (*F_o*² + 2*F_c*²)/3

Other non-hydrogen atoms were found from successive difference Fourier synthesis. Anisotropic thermal parameters for all non-hydrogen atoms were included in the refinements. All hydrogen atoms bonded to carbon atoms were included in calculated positions. This C-H bond distance was fixed and U values were assigned based approximately on the U value of the attached atom. The other hydrogens (N-H) were included in located positions with $U = 0.05 \text{ \AA}^2$. A final difference Fourier map revealed several random features ($<0.56 \text{ e\AA}^{-3}$). Crystal and refinement data of the complex are summarized in Table 1.

Results and Discussion

Synthesis and Characterization. The only product prepared from the reaction of L^1 with excess acrylonitrile in methanol was the di-*N*-cyanoethylated macrocycle L^2 . Such a di-*N*-cyanoethylation has been also observed for the reaction of L^7 with excess acrylonitrile.¹³ However, present result is different from the fact that the tetra-*N*-cyanoethylated macrocycle L^{10} as well as L^{11} can be prepared readily by the reaction of L^9 with acrylonitrile under the similar condition.¹⁰ The macrocycle L^2 dissolves in chloroform but slightly in methanol at room temperature. The mass, infrared, and ¹³C-NMR spectral data of the compound are listed in the Experimental. The position of the cyanoethyl pendant arms was confirmed by the crystal structure of $[\text{NiL}^2](\text{ClO}_4)_2$ (see below).

The reaction of L^8 with $\text{Ni}(\text{OAc})_2 \cdot 4\text{H}_2\text{O}$ followed by addition of NaClO_4 produced the square-planar complex $[\text{NiL}^8](\text{ClO}_4)_2$. The similar reaction of L^2 with $\text{Ni}(\text{OAc})_2 \cdot 4\text{H}_2\text{O}$ produced only the green complex $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$, in which the acetate ion is coordinated to the metal ion. The square-planar complex $[\text{NiL}^2](\text{ClO}_4)_2$ can be prepared by adding HClO_4 to a hot aqueous solution of $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$. As described in Experimental, however, the reaction conditions such as pH and temperature must be carefully controlled for the preparation of the square-planar complex. For example, addition of HClO_4 to the solution of $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$ at room temperature produces the white salt $[\text{H}_2\text{L}^2](\text{ClO}_4)_2$. It should be noted that the complex $[\text{NiL}^2](\text{ClO}_4)_2$ could not be prepared by direct reaction of L^2 with Ni^{2+} ion. Addition of NaClO_4 to a methanol solution (reflux >3 h) of the macrocycle and hydrated or dehydrated NiX_2 ($\text{X} = \text{CH}_3\text{COO}^-$, Cl^- , or NO_3^-) produced only $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$ or $[\text{H}_2\text{L}^2](\text{ClO}_4)_2$. Present result is quite different from those reported for most other di-*N*-substituted 14-membered tetraaza macrocycles;^{7,9,10,19} the square-planar complexes $[\text{NiL}](\text{ClO}_4)_2$ ($\text{L} = \text{L}^3, \text{L}^4, \text{L}^8$, or L^{11}) can be prepared readily by direct reaction of each macrocycle with hydrated or dehydrated Ni^{2+} ion.

The complexes $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$ and $[\text{NiL}^2](\text{ClO}_4)_2$ are soluble in nitromethane and dimethylsulfoxide. Infrared, ¹³C-NMR, and/or FAB-mass spectra of the complexes are summarized in Experimental, together with elemental analyses. The molar conductance values of $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$ ($85 \text{ \Omega}^{-1}\text{mol}^{-1}\text{cm}^2$) and $[\text{NiL}^2](\text{ClO}_4)_2$ ($140 \text{ \Omega}^{-1}\text{mol}^{-1}\text{cm}^2$)

Table 2. Electronic Absorption Spectral Data^a

Complex	λ_{max} , nm (ϵ , $\text{M}^{-1}\text{cm}^{-1}$)
$[\text{NiL}^1](\text{ClO}_4)_2$ ^b	469(83)
$[\text{NiL}^2(\text{OAc})]\text{ClO}_4$	406(61) 705(24) 405(50) ^f 695(20) ^f
$[\text{NiL}^3](\text{ClO}_4)_2$	515(150) 518(145) ^f
$[\text{NiL}^3](\text{ClO}_4)_2$ ^b	505(120) 506(118) ^f
$[\text{NiL}^4](\text{ClO}_4)_2$ ^b	502(110) 502(112) ^f
$[\text{NiL}^8](\text{ClO}_4)_2$	496(120) 495(115) ^f
$[\text{NiL}^{11}](\text{ClO}_4)_2$ ^e	490(117) 490(115) ^f

^aIn nitromethane at 20 °C unless otherwise specified. ^bRef. 9. ^cIn acetonitrile. ^dIn dimethylsulfoxide. ^eRef. 10.

measured in nitromethane indicate that the complexes are 1 : 1 and 1 : 2 electrolytes, respectively.

In the electronic absorption spectra (Table 2), the *d-d* band of $[\text{NiL}^2](\text{ClO}_4)_2$ is observed at much longer wavelength than that of $[\text{NiL}^1](\text{ClO}_4)_2$, as usual. Interestingly, the wavelengths of the di-*N*-cyanoethylated macrocyclic complexes are strongly affected by the number or type of the *C*-substituents and show the order L^{11} (490 nm) $< L^8$ (*ca.* 495) $\ll L^2$ (*ca.* 515 nm). Furthermore, the wavelength of $[\text{NiL}^2](\text{ClO}_4)_2$ is also *ca.* 10 nm longer than those for the complexes of L^3 and L^4 containing two *N*-alkyl groups. To our knowledge, $[\text{NiL}^2](\text{ClO}_4)_2$ is a rarely prepared di-*N*-substituted 14-membered tetraaza macrocyclic nickel(II) complex that show the *d-d* band at such a long wavelength (>510 nm). The weaker ligand strength of $[\text{NiL}^2](\text{ClO}_4)_2$, compared to that of $[\text{NiL}](\text{ClO}_4)_2$ ($\text{L} = \text{L}^3, \text{L}^4, \text{L}^8$, or L^{11}), may be attributed to the weaker Ni-N interactions due to the severe steric crowding caused by the presence of the eight *C*-methyl groups together with the *N*-cyanoethyl pendant arms.^{3-5,7-10} The spectra of $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$ measured in nitromethane or acetonitrile show two bands at *ca.* 405 and 700 nm, which are comparable with those of various five-coordinate nickel(II) complexes.²⁰⁻²² Unfortunately, we could not grow single crystals of $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$ suitable for the X-ray work, and coordination geometry of the complex is not clearly understood at this point.

Crystal Structure of $[\text{NiL}^2](\text{ClO}_4)_2$. The structure (Figure 1) of $[\text{NiL}^2]^{2+}$ shows that the complex has a square-planar coordination geometry with an inversion center at the metal atom. The cyanoethyl groups are not involved in the coordination. The six-membered chelate rings have a chair conformation, and the five-membered chelate rings a *gauche* one. The macrocyclic ligand adopts the *trans*-III conformation.

Selected bond distances and angles of the complex are listed in Table 3. The N-Ni-N bond angle ($86.6(2)^\circ$) involved in the five-membered chelate ring is smaller than that ($93.4(2)^\circ$) in the six-membered chelate ring, as usual. The C(10)-N(2)-Ni angle is $100.4(2)^\circ$. The C(4)-N(2)-C(10) angle ($114.0(2)^\circ$) is larger than the C(3)-N(2)-C(10) angle. The N(3)-C(12) bond distance ($1.131(5) \text{ \AA}$) is corresponding to the triple bond. The Ni-N(1) (secondary) and Ni-N(2) (tertiary) distances ($1.972(2)$ and $2.008(2) \text{ \AA}$, respectively) are comparable with those of the square-planar complex

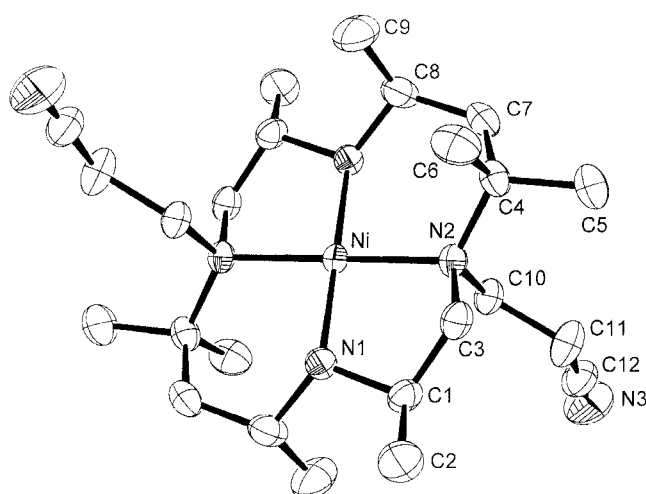


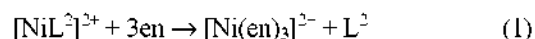
Figure 1. An ORTEP diagram of $[\text{NiL}^2]^{2+}$ in $[\text{NiL}^2](\text{ClO}_4)_2$.

$[\text{NiL}^3](\text{ClO}_4)_2$ (1.984(2) and 2.000(2) Å)⁹ but are distinctly longer than those of other related nickel(II) complexes with di-*N*-alkylated 14-membered tetraaza macrocycles.^{5,7} One of the interesting structural features of the complex is that the Ni-N(2) distance is much longer than the Ni-N(1) distance. The N(2)-C(4) distance (1.549(3) Å) is also considerably longer than other N-C single bond distances. The relatively long Ni-N(2) and N(2)-C(4) distances may be resulted from the severe steric crowding caused by both the methyl groups at the C(4) atom and the *N*-cyanoethyl groups. The crystallographic study shows that the unusually weak ligand field of the complex is correlated with the long Ni-N(2) distance.

Solution Behaviors. The complex $[\text{NiL}^2](\text{ClO}_4)_2$ is relatively stable in neutral and acidic aqueous solutions at room temperature. Electronic absorption spectra of the complex (3.0×10^{-3} M) measured in 0.1 M HClO_4 aqueous solution indicated that no apparent decomposition of the complex or decyanoethylation takes place in 20 h at 25 °C. On the other hand, the complex was completely decyanoethylated to $[\text{NiL}^1]^{2-}$ within 3 h in boiling aqueous solution. Such behaviors are not so much different from those observed for $[\text{NiL}^8]^{2-}$ and $[\text{NiL}^{11}]^{2+}$ in neutral or acidic aqueous solutions.^{10,19}

In a DMSO-water (1 : 1) solution containing NaOAc (0.1

M), $[\text{NiL}^2](\text{ClO}_4)_2$ rapidly reacts with acetate ion to form the green complex $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$ even at room temperature. This was confirmed by the observation that electronic absorption spectrum of the solution was essentially the same as that of $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$. Addition of NaClO_4 to the solution also produced the green complex. On the other hand, addition of a slightly excess amount of ethylenediamine (en) to a DMSO or DMSO-water (1 : 1) solution of $[\text{NiL}^2](\text{ClO}_4)_2$ produced L^2 and $[\text{Ni}(\text{en})_3](\text{ClO}_4)_2$, indicating that the macrocyclic ligand is substituted by en. The complex $[\text{NiL}^2](\text{ClO}_4)_2$ was decomposed rapidly to yield L^2 in a DMSO-water (1 : 1) solution containing NaOH (*ca.* 0.01 M). The substitution and addition reactions (Eqs. (1) and (2), respectively) were too fast to measure with ordinary spectroscopic methods even at room temperature. Interestingly, such behaviors are in sharp contrast to the generally observed trend that most square-planar nickel(II) complexes of 14-membered tetraaza macrocycles are quite stable and are hardly attacked by en or acetate ion at room temperature.^{10,19,23} Although $[\text{NiL}^3]^{2+}$ and $[\text{NiL}^4]^{2+}$ are decomposed in NaOH solutions,⁸ they do not react with acetate ion even in >1.0 M NaOAc solutions. Present result is also different from the observation that $[\text{NiL}^8]^{2+}$ and $[\text{NiL}^{11}]^{2-}$ are not decomposed even at 0.1 M NaOH solutions but are rapidly decyanoethylated to $[\text{NiL}^8]^{2-}$ or $[\text{NiL}^9]^{2-}$.^{10,19} The enhanced lability of $[\text{NiL}^2](\text{ClO}_4)_2$ in the solutions of NaOAc, en, NaOH can be attributed to the relatively long Ni-N(2) distance.



In general, both the electronic and steric effects are important factors affecting the Lewis basicity and coordination behavior of alkyl amines. However, the relatively weak Ni-N interaction and the unusual reactivity of $[\text{NiL}^2]^{2-}$ correspond to the suggestion that the electronic effects are exceeded by the back strain of the tertiary amino groups caused by the steric repulsion between the *N*-cyanoethyl pendant arms and proximate C-methyl groups.⁷⁻¹⁰

Conclusion

The di-*N*-cyanoethylated macrocycle L^2 and its nickel(II) complexes $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$ and $[\text{NiL}^2](\text{ClO}_4)_2$ were prepared. The square-planar complex $[\text{NiL}^2]^{2+}$, which can be prepared by addition of HClO_4 to a hot solution of $[\text{NiL}^2(\text{OAc})]\text{ClO}_4$, exhibits unusually weak ligand field strength and rapidly reacts with acetate ion or en even at room temperature. The synthetic procedure, ligand field strength, and reactivity of $[\text{NiL}^2]^{2-}$ are distinctly different from those of $[\text{NiL}^3]^{2+}$, $[\text{NiL}^8]^{2+}$, or $[\text{NiL}^{11}]^{2+}$, showing that they are strongly affected by the steric repulsion between *N*-cyanoethyl groups and proximate C-methyl groups.

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Table 3. Selected Bond Distances (Å) and angles (deg) of $[\text{NiL}^2](\text{ClO}_4)_2$

Ni-N(1)	1.972(2)	Ni-N(2)	2.008(2)
N(1)-C(1)	1.502(3)	N(2)-C(3)	1.494(3)
N(2)-C(4)	1.549(3)	N(2)-C(10)	1.511(3)
N(3)-C(12)	1.131(5)	C(1)-C(3)	1.487(4)
C(10)-C(11)	1.536(4)	C(11)-C(12)	1.452(5)
N(1)-Ni-N(2)	86.6(2)	N(1)-Ni-N(2)	93.4(2)
N(3)-C(12)-C(11)	179.2(4)	C(10)-C(11)-C(12)	108.9(3)
C(4)-N(2)-Ni	117.7(2)	C(1)-N(1)-Ni	110.1(2)
C(10)-N(2)-Ni	100.4(2)	C(3)-N(2)-C(10)	111.2(2)
C(4)-N(2)-C(10)	114.0(2)		

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