## Notes

# Determination of Viscosity Average Molecular Weight from Number and Weight Average Molecular Weights

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**Abstract**: A theoretical equation for the viscosity average molecular weight is derived as  $\overline{M}_v = \{(1/y^a), (\Gamma(z+a+1)/(\Gamma(z+1))\}^{1/a}(M_0), \text{ where the adjustable parameters, } y \text{ and } z, \text{ of the Schultz-Zimm distribution can be determined from the number and weight average molecular weights. When the value of the exponent, <math>a$ , appearing in the relation between the intrinsic viscosity and the viscosity average molecular weight, is in the range from 0.6 to 0.8, as is the case for the most polymers, the viscosity average molecular weight is much more close to the weight average molecular weight than to the number average molecular weight.

#### Introduction

The number and weight average degrees of polymerization can be determined theoretically from polymerization mechanism and kinetics when reaction rate constants of the elementary reactions are constant as is the case with polymerizations in continuous stirred tank reactors. However, many industries use the viscosity average molecular weight in place of the number and weight average molecular weights for their product control.

In this study, a theoretical equation is derived to correlate the viscosity average molecular weight with the number and weight average molecular weights when the molecular weight distribution of the polymers follows the Schultz-Zimm distribution.

#### Theoretical Equations

The average intrinsic viscosity is given as a function of viscosity average molecular weight as Eq(1).

$$[\overline{\eta}] = K \overline{M}_v^a \tag{1}$$

the average intrinsic viscosity becomes as in Eq(3)

$$[\eta]_i = KM_i^a \tag{2}$$

For a molecule whose molecular weight is  $M_i$ ,

the intrinsic viscosity is expressed as Eq(2), and

$$\overline{[\eta]} = w_i \ [\eta_i] = K \sum w_i M_i^a$$
 (3)

Therefore, viscosity average molecular weight yields to be Eq(4).

$$\overline{M_v} = \left\{ \sum w_i M_i^a \right\}^{1/a} = \left\{ \sum (i^a w_i) \right\}^{1/a} M_0$$
 (4)

where  $w_i$  is weight fraction of molecules whose degree of polymerization is i, and  $M_0$  is average molecular weight of the structural unit constituting the polymer molecule.

Because the weight fraction  $w_i$  can be expressed as Eq(5), the rearrangement of Eq(4) yields Eq(6)

$$w_i = \frac{iP_i}{\sum iP_i} \tag{5}$$

$$M_{\nu} = \left\{ \frac{\sum_{i} i^{a+1} P_{i}}{\sum_{i} i P_{i}} \right\}^{\frac{1}{a}} M_{0} \tag{6}$$

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where  $P_i$  is the concentration of polymer molecules whose degree of polymerization is i.

The Schultz-Zimm distribution function for weight fraction,  $w_i$ , and for number fraction,  $n_i$ , can be written as Eq(7) and (8).<sup>3</sup>

$$w_{i} = \frac{y^{z+1}}{\Gamma(z+1)} i^{z} e^{-yi} \tag{7}$$

$$n_i = \frac{y^z}{\Gamma(z)} i^{z-1} e^{-yi} \tag{8}$$

Therefore, the weight-average and number-average degree of polymerization  $\mu_w$  and  $\mu_n$ , can be derived to be as Eq(9) and (10).<sup>2</sup>

$$\mu_w = \int_0^\infty i w_i di = \frac{z+1}{v} \tag{9}$$

$$\mu_n = \int_0^\infty i n_i di = \frac{z}{v} \tag{10}$$

where both z and y can be determined from  $\mu_w$  and  $\mu_n$ .

From Eq(4) and (7) the viscosity-average degree of polymerization,  $\mu_{\nu}$  can be shown as

$$\mu_{v} = \left\{ \frac{v^{z+1}}{\Gamma(z+1)} \int_{0}^{\infty} i^{z+a} e^{-yi} di \right\}^{\frac{1}{a}}$$
 (11)

Hence,  $\mu_v$  and  $\overline{M}_v$  can be calculated by Eq(12) and Eq(13), respectively, from the adjustable parameters, y and z, of the Schulz-Zimm distribution function.

$$\mu_{v} = \left\{ \frac{1}{v^{a}} \frac{\Gamma(z+a+1)}{\Gamma(z+1)} \right\}^{\frac{1}{a}}$$
 (12)<sup>4</sup>

$$\overline{M}_{v} = \mu_{v} M_{0} \tag{13}$$

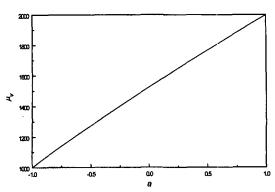
## **Discussion**

Randomly coiled polymers have the value of the exponent, a, in the range between 0.5 and 1.0

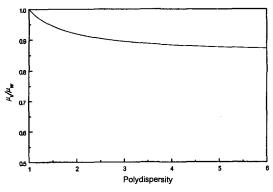
Figure 1 shows  $\mu_v$  as a function of a for  $\mu_n = 1,000$  and  $\mu_w = 2,000$ . When a is equal to -1.0,  $\mu_v$  is equal to  $\mu_n$ . As a approaches to 1.0,  $\mu_v$ 

increases almost linearly to  $\mu_w$ . Since most polymers have their a values between 0.6 and 0.8,  $\mu_v$  is more inclined to  $\mu_w$  than to  $\mu_n$ .

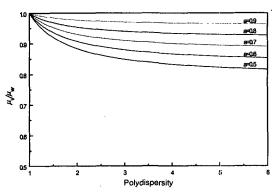
Figure 2 demonstrates the variation of  $\mu_v/\mu_w$  as



**Figure 1.**  $\mu_v$  as a function of the exponent a for a polymer whose  $\mu_n$  and  $\mu_w$  are 1,000 and 2,000, respectively.



**Figure 2.** Variation of  $\mu_v/\mu_w$  as a function of the polydispersity,  $\mu_w/\mu_n$ . The value of the exponent, a, was fixed at 0.65. The curves for different values of  $\mu_w$  overlapped completely with each other.



**Figure 3.** Dependence of  $\mu_w/\mu_w$  on  $\mu_w/\mu_n$  the polydispersity index at different values of the exponent, a.

Determination of Viscosity Average Molecular Weight From Number and Weight Average Molecular Weights

a function of  $\mu_{\omega}/\mu_{n}$ , the polydispersity index, at a fixed value of a. The curves for different values of  $\mu_{\omega}$ , overlap with each other completely, indicating that dependence of the ratio,  $\mu_{\omega}/\mu_{\omega}$ , on the polydispersity index is independent of the value of  $\mu_{\omega}$ .

Dependence of  $\mu_v/\mu_w$  on the polydispersity index at different values of a is shown in Figure 3.  $\mu_v$  deviates from  $\mu_w$  as the polydispersity index is increased. However, it can be perceived that  $\mu_v$  is not too much smaller than  $\mu_w$  for most polymers.

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### References

- (1) M. Kurata, Y. Tsunashima, M. Iwama, and K. Kamada, Polymer Handbook, J. Brandrup and E. H. Immergut, Eds., John Wiley and Sons, New York, 1975, Ch. 4.
- (2) F. W. Billmeyer, Text book of Polymer Science, 3rd Ed., John Wiley and Sons, New York, 1984, Ch. 8.
- (3) B. H. Zimm, J. Chem. Phys., 16, 1099 (1948).
- (4) A computer program is available upon request for calculation of the gamma function.