

Fabrication of Cordierite Honeycomb from Fly Ash

Sung-Jin Kim^{1,a}, Sung-Jin Park^{1,b}, Hee-Gon Bang^{1,2,c}, Sang-Yeup Park^{1,2,d}

¹Department of Ceramic Engineering, Kangnung National University,
Kangnung, Kangwondo, 210-702, South Korea

²Technology Innovation Center for Fine Ceramics, Kangnung National University,
Kangnung, Kangwondo, 210-702, South Korea

^ak9s8j7@hotmail.com, ^bdark sung jin@hanmail.net,

^chgbang@kangnung.ac.kr ^dsypark@knusun.kangnung.ac.kr

Abstract

In this study, we attempt to synthesize the cordierite from the reaction of fly-ash, alumina, silicon dioxide, and magnesia powders. For the purpose of air purification, the honeycomb filter with porous cordierite was fabricated from the combination of synthetic cordierite and pore forming agent. Fabricated porous cordierite honeycomb was prepared with high porosity (58%), and good compressive strength (69MPa).

Keywords : Cordierite, Fly ash, Porous honeycomb filter, Extrusion process.

1. Introduction

Cordierite honeycomb with porous body has a great advantage in the industrial applications such as catalyst carrier for automobile and high temperature gas filter [1-4]. Key factors in the application of high temperature filter are known as high surface area and compressive strength. Aluminium titanate, mullite, zirconia, spodumene, and cordierites with high surface area and good adsorption properties are used as a ceramic filter [5]. Among those materials, cordierite has been a good candidate due to its high thermal shock resistance and low thermal expansion coefficient [6, 7]. Because the major chemical components of cordierite are Al₂O₃, SiO₂, MgO, cordierite is fabricated by fly ash, talc, and kaolin as a raw materials. In this study, we attempted to fabricate the cordierite honeycomb using fly ash known as an industrial waste. Also the effect of graphite additive as a porosifier as well as lubricant during the extrusion process of honeycomb body was investigated.

2. Experimental and Results

As a starting powder for cordierite synthesis, fly ash (Youngdong power plant, Korea), MgO (Kosundo chemical, Japan) were used. Chemical composition of fly ash was shown in table 1. Cordierite powder was synthesized from the solid state reaction method at 1100°C. Synthesized cordierite was mixed with binder (Methyl Cellulose, Showa Chemical, Japan), plasticizer, (Glycerol, Showa Chemical, Japan), and lubricant (LU-6418, Cell Chemical, Korea) for the extrusion of honeycomb body. During the extrusion of cordierite honeycomb, extrusion pressure was controlled. After extrusion, honeycomb was dried for 12 h with the

control of moisture and temperature, and then heat treated at 1000°C ~ 1250°C. Extrusion additives (binder, plasticizer, lubricant) were varied to control the optimum condition for honeycomb properties. Also, graphite was added to replace the lubricant as well as to enhance the pore formation. Table. 2 shows the compositions of honeycomb used in this study. Physical properties of honeycomb were investigated using porosimeter, compressive stress (Instron, 5882, USA).

Table. 1 Chemical composition of fly ash.

Component	Weight Percent (wt%)								
	SiO ₂	Al ₂ O ₃	Fe ₂ O ₃	K ₂ O	TiO ₂	CaO	SO ₃	MgO	Na ₂ O
Fly Ash	48.1	32.9	6.43	5.21	2.28	1.54	1.21	0.970	0.421

Table. 2 Specimen notations of cordierite honeycomb.

Notation	Composition
C-17A	Cordierite + 17wt% Additives
C-17A10G	Cordierite + 17wt% Additives + 10wt% Graphite
C-17A30G	Cordierite + 17wt% Additives + 30wt% Graphite
C-34A	Cordierite + 34wt% Additives
C-34A10G	Cordierite + 34wt% Additives + 10wt% Graphite

Fig. 1 shows the variation of honeycomb shrinkage with the extrusion additives. This results suggest that the impurity content in graphite powder was affected the densification of cordierite honeycomb due to the liquid

phase formation. This prediction was confirmed from rapid shrinkage rate above 1200°C due to the particle rearrangement and the viscous flow onset of liquid formation.

Fig. 2 shows the variation of extrusion pressure with extrusion additives as well as graphite addition. All the extrusion condition, extrusion rate was increased with increasing the extrusion pressure. However, extrusion pressure was decreased with increasing the graphite addition. This result implied that graphite addition was resulted from the enhancement of flow ability due to the relaxation of fraction stress between cordierite particles.

Fig. 3 shows the variation of porosity and compressive strength with sintering temperature and extrusion composition. With increasing the extrusion additives, porosity was increased. However, graphite addition was not effective for the porosity increase. This result was confirmed in the compressive strength data. Therefore, graphite addition was effective only for the enhancement of extrusion rate similar to lubricant additive due to the impurity content in graphite powder.

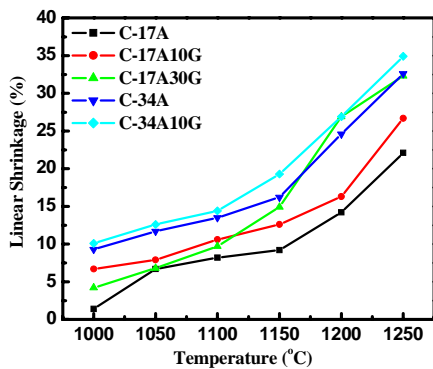


Fig. 1. Linear shrinkage of cordierite with different amounts of organic matter and graphite.

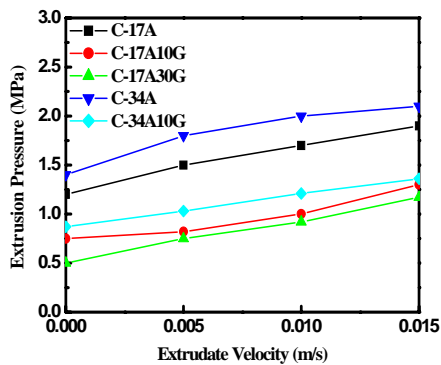


Fig. 2. Effect of organic matter and graphite addition on extrusion pressure.

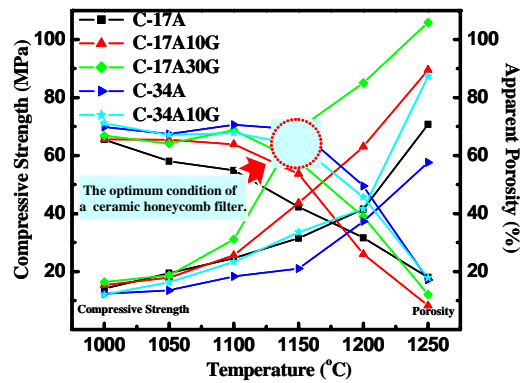


Fig. 3. Variation of compressive strength and apparent porosity in honeycomb structure with different amounts of organic matter and graphite.

3. Summary

Cordierite honeycomb filters were fabricated using fly ash. With increasing organic binder content, porosity in the honeycomb increased and compressive strength was decreased. Porous cordierite honeycomb with high porosity (58%) and good compressive strength (69MPa) was obtained by the addition of 30wt% graphite at low sintering temperature (1150°C). Although graphite addition was not effective for the increase of porosity in honeycomb body, it was effective for the increase of compressive strength and the decrease of extrusion pressure during the honeycomb process. This result is believed that the impurity content in the graphite was attributed to the enhancement of densification and particle binding force in the ceramic honeycomb body.

4. References

1. Jin Yang: Journal of the Korean Ceramic Society Vol.35, No. 4, pp. 399-405 (1998).
2. J. D. Chung, Environ. Eng. Res., 2(1), 33-40 (1997).
3. A. Bueno-Lopez, D. Lozano-Castello, I. Such-Basanez, J. M. Garcia-Cortes, M. J. Illan-Gomez, C. Salinase-Martinez de Lecea, Applied Catalysis B : Environmental 58, 1-7 (2005).
4. Alessandro Bachiarrini, Ceramics International 22, 73-77 (1996).
5. Y. Sawada, K. Hiramatsu, H. Kawamoto et al: High Temperature Gas Cleaning, Volume II, 393-404 (1999).
6. Young-Bae Kim, Eul-Hoon Cho, Yoon-Young Chang, Hee-Soo Lee and Duck-Kyun Choi: Korean Journal of Material Research Vol. 12, No. 2 (2002).
7. M.A.Camerucci, G.Urretavizcaya, A.L.Cavaliere : Ceramics International 29 (2003) 159-168.