

Density Functional Theory Study of UO_2

Younsuk Yun,^aHanchul Kim,^bHeemoon Kim,^cKwangheon Park ^a

^a Department of Nuclear Engineering, Kyung-Hee University, Suwon 449-701, Korea

^b Korea Research Institute of Standard and Science, P.O. Box 102, Yuseong, Daejeon 305-600 Korea

^c Korea Atomic Energy Research Institute, P.O. Box 105, Yuseong, Daejeon 305-353 Korea

1. Introduction

Uranium dioxide (UO_2) is an important fuel material for nuclear industry, but the electronic structure of $5f$ electrons of uranium has not been reliably described so far. The interpretation of some of experimental results is still controversial because of a large variety of unusual electronic effects originating from strong Coulomb repulsion between $5f$ electrons of uranium. During irradiation, UO_2 undergo nuclear fission and the noble gas Xe is formed. Xe diffuses into the gap between the cladding and pellets of the fuel, and causes swelling of the fuel. Point defects are thought to be the major channel of Xe diffusion in UO_2 [ⁱ, ⁱⁱ]. Therefore, it is important to investigate defects in order to understand the diffusion mechanism in nuclear reactor. We present the density functional theory (DFT) calculations on UO_2 employing the local spin density approximation (LSDA) + U method. First, we calculated the chemical and physical properties of bulk UO_2 such as the lattice constant, bulk modulus, cohesive energy, magnetic moment, and band gap. We compared the calculated results with available experimented data. Then we performed supercell calculations to investigate the structure of different defects in UO_2 , the defect formation energy, and the defect-induced changes in the electronic structure.

2. Theory and methodology

In recent years, DFT calculations have made a profound impact on the investigation of material properties. In addition, improvements in computer performance allow to apply these methods to an increasing number of

physical and chemical phenomena. Most of DFT calculations are based on the local density approximation (LDA) of the exchange-correlation [ⁱⁱⁱ, ^{iv}]. However the LDA calculations often fail to describe systems with strongly correlated $5f$ electrons, predicting metallic behaviour contrary to the observed insulating behaviour [^v]. Orbital-dependent functionals like the LSDA+U approach are known to correct this problem by adding on-site Coulomb repulsion (\bar{U}). The $5f$ Coulomb correlation energy \bar{U} was determined to be $4.6 \pm 0.8 \text{ eV}$ from the energy difference in the two spectroscopy experiments (x-ray photoemission and Bremsstrahlung isochromat spectroscopy) [^{vi}]. We used the parameters of $\bar{U} = 4.5 \text{ eV}$ and $\bar{J} = 0.5$ from Dudarev *et al.* [^{vii}].

3. Results and discussion

Table. 1. shows the calculated chemical and physical properties of UO_2 . The calculated equilibrium lattice constant of UO_2 is found to be 5.44 \AA . It is underestimated by about 0.55 % compared with the experimental value of 5.47 \AA [^{viii}]. The cohesive energy per UO_2 molecule is calculated to be 26.9 eV , that is larger than the experimental value of 22.3 eV by about 20 %. The underestimation of lattice constant and the overestimation of cohesive energy are typical of DFT-LDA calculations [^{ix}]. The LSDA+U calculations predict the correct insulating ground states, which shows antiferromagnetic ordering, with the band gap of 1.9 eV . In contrast, calculations employing original LDA scheme result in metallic electronic structure with the partially occupied uranium $5f$ bands.

Table. 1 Cohesive properties of UO_2

	Exp.	S.L. Dudarev (LSDA)	S.L. Dudarev (LMTO-ASA)	This work (LSDA+U)
Lattice constant (\AA)	5.46	5.27	5.37	5.44
Cohesive energy (eV/ UO_2)	22.31	18.63	22.23	26.90*
Bulk modulus (GPa)	207	252	202	209
Magnetic moment (μ_B)	1.74	–	1.7	1.9
Band gap. (eV)	2.0	–	1.3	1.9

* We did not include spin-polarization energy of atoms. Thus our calculated cohesive energy is somewhat overestimated.

REFERENCES

-
- [ⁱ] J. P. Crocombette, F. Jollet, L.Thien Nga, and T.Petit, Phys. Rev. B **64**, 104107
- [ⁱⁱ] Richard G. J. Ball, Robin W. Grimes, J. Chem. Soc. Faraday Trans., 1990, **86**(8), 1257-1261
- [ⁱⁱⁱ] D.R. Hanmann, M. Schluter, and C. Chiang, Phys. Rev. Lett. **43**, 1494(1979)
- [^{iv}] P. Hohenberg and W.Kohn, Phys. Rev. **136**, B664 (1964)
- [^v] G. A. Sawatzky and J. W. Allen, Phys. Rev. Lett. **53**, 2339 (1984)
- [^{vi}] Y. Baer and J. Schoenes, Phys. Rev. B **47**, 888 (1980)
- [^{vii}] S. L. Dudarev, G. A. Botton, S. Y. Savrasov, C. J. Humphreys, A. P. Sutton, Phys. Rev. B **57**(3), 1505(1998).
- [^{viii}] L. Lynds, J.Inorg. Nucl. Chem. **24**, 1007 (1962)
- [^{ix}] Paxton, A. T., Methfessel, M., and Polatoglou, H. M., 1990. Phys. Rev. B **41**, 8127